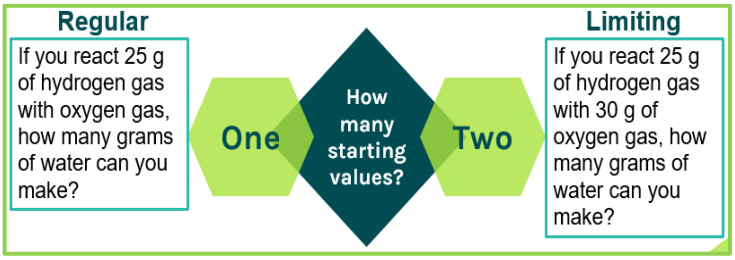
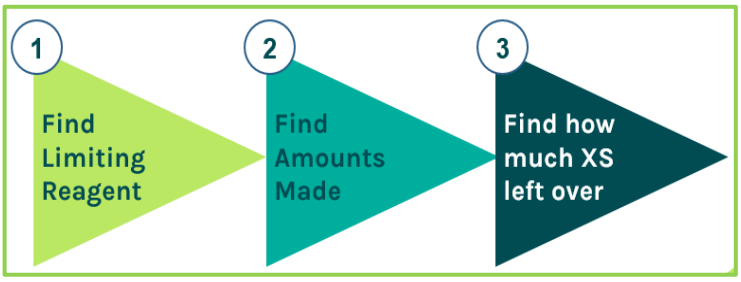
Regular stoich or limiting stoich?



Three main types of problems:



Steps

1. Grams to moles
2. Have vs. need
3. Identify limiting
4. Stoich with limiting *(if asked)*
5. Find xs left *(if asked)*

Practice Problem #1:

If you reacted 150.0 g of K with 225 g of Br_2, how may g of KBr can be made? How much excess reagent is left?

$2K + Br_2 \rightarrow 2KBr$

Practice Problem #2:

If you react 13.2 g of Fe with 6.34 g of O_2, how may g of Fe_2O_3 are made? How many grams of excess are left?

$4Fe + 3O_2 \rightarrow 2Fe_2O_3$

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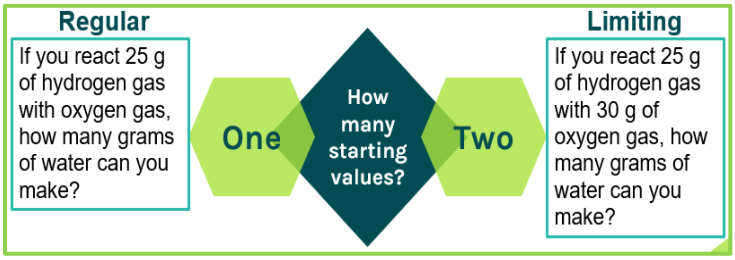

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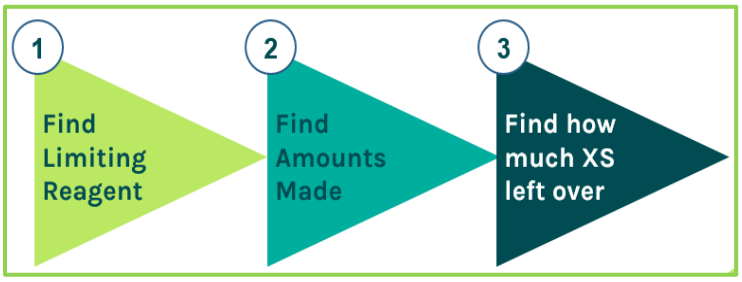

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